

#### Classical Dynamics at 300K

Energy function:  $U(\vec{r}_1, \vec{r}_2, \cdots, \vec{r}_N) = U(\vec{R})$ 

used to determine the force on each atom:

$$m_i \frac{d^2 \vec{r_i}}{dt^2} = \vec{F_i} = -\vec{\nabla} U(\vec{R})$$

yields a set of 3N coupled 2<sup>nd</sup>-order differential equations that can be propagated forward (or backward) in time.

Initial coordinates obtained from crystal structure, velocities taken at random from Boltzmann distribution.

Maintain appropriate temperature by adjusting velocities.







## Molecular Dynamics Ensembles

Constant energy, constant number of particles (NE)

Constant energy, constant volume (NVE)

Constant temperature, constant volume (NVT)

Constant temperature, constant pressure (NPT)

Choose the ensemble that best fits your system and start the simulations, but use NE to check on accuracy of the simulation.

# Langevin Dynamics for temperature control Langevin dynamics deals with each atom separately, balancing a small friction term with Gaussian noise to control temperature: $m \, \ddot{\vec{r}} = \vec{F}(\vec{r}) - \gamma \, m \, \dot{\vec{r}} + \vec{R}(t)$ $\langle \vec{R}(t) \cdot \vec{R}(t') \rangle = 6k_B T \gamma \, \delta(t - t')$



## Large is no problem. But ...





## Preparing Your System for MD Solvation

Biological activity is the result of interactions between molecules and occurs at the interfaces between molecules (protein-protein, protein-DNA, protein-solvent, DNA-solvent, etc).

Why model solvation?

• many biological processes occur in aqueous solution

• solvation effects play a crucial role in determining molecular conformation, electronic properties, binding energies, etc

How to model solvation?

• explicit treatment: solvent molecules are added to the molecular system

• implicit treatment: solvent is modeled as a continuum dielectric or so-called implicit force field

*mitochondrial bc1 complex* 



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(Usually periodic! Avoids surface effects)





